Dublin Institute of Technology
ARROW@DIT

Articles

Crest: Centre for Research in Engineering Surface Technology

2013-02-01

# Anti-Bacterial Activity of Indoor Light Activated Photocatalysts

Suresh Pillai
*Dublin Institute of Technology*, suresh.pillai@dit.ie

Michael K. Seery
*Dublin Institute of Technology*, michael.seery@dit.ie

Damian W. Synnott
*Dublin Institute of Technology*

Stephen Hinder
*Univerity of Surrey*

Georg Michlits
georg.michlits@gmail.com

Follow this and additional works at: http://arrow.dit.ie/cenresart

Part of the Nanoscience and Nanotechnology Commons

Recommended Citation
Synnott, D. Seery, M. Hinder, S. Michlitsand, G. Pillai, S. Anti-bacterial activity of indoor light activated photocatalysts, Applied Catalysis B, 130– 131, 2013, 106– 111

This Article is brought to you for free and open access by the Crest: Centre for Research in Engineering Surface Technology at ARROW@DIT. It has been accepted for inclusion in Articles by an authorized administrator of ARROW@DIT. For more information, please contact yvonne.desmond@dit.ie, arrow.admin@dit.ie.

This work is licensed under a Creative Commons Attribution-Noncommercial-Share Alike 3.0 License

Accepted Manuscript

Title: Anti-Bacterial Activity of Indoor-Light Activated Photocatalysts

Authors: Damian W. Synnott, Michael K. Seery, Steven J. Hinder, Georg Michlits, Suresh C. Pillai

PII: S0926-3373(12)00488-2
DOI: doi:10.1016/j.apcatb.2012.10.020
Reference: APCATB 12301

To appear in: *Applied Catalysis B: Environmental*

Received date: 29-5-2012
Revised date: 16-10-2012
Accepted date: 22-10-2012

Please cite this article as: D.W. Synnott, M.K. Seery, S.J. Hinder, G. Michlits, S.C. Pillai, Anti-Bacterial Activity of Indoor-Light Activated Photocatalysts, *Applied Catalysis B, Environmental* (2010), doi:10.1016/j.apcatb.2012.10.020

This is a PDF file of an unedited manuscript that has been accepted for publication. As a service to our customers we are providing this early version of the manuscript. The manuscript will undergo copyediting, typesetting, and review of the resulting proof before it is published in its final form. Please note that during the production process errors may be discovered which could affect the content, and all legal disclaimers that apply to the journal pertain.

*Manuscript
ACCEPTED MANUSCRIPT

# Anti-Bacterial Activity of Indoor-Light Activated Photocatalysts

Damian W. Synnott[a,b], Michael K. Seery[b], Steven J. Hinder[c], Georg Michlits[b] and Suresh C. Pillai[a*]

[a] Centre for Research in Engineering Surface Technology (CREST), FOCAS Institute, Dublin Institute of Technology, Camden Row, Dublin 8, Ireland

[b] School of Chemical and Pharmaceutical Sciences, Dublin Institute of Technology, Kevin St. Dublin 8, Ireland

[c] The Surface Analysis Laboratory, Faculty of Engineering and Physical Sciences, University of Surrey, GU2 7XH, United Kingdom

Corresponding authors email: suresh.pillai@dit.ie

**Abstract**

Nanocrystalline photocatalysts, prepared under ambient conditions using a microwave assisted synthesis, show indoor light photocatalytic activity for the degradation of *S. aureus* and *E. coli*. The zinc sulfide (ZnS) nanomaterials, prepared by a microwave assisted synthesis, are shown to be cubic blende structure with an average crystallite size of 4 – 6 nm. The anti-bacterial activity of these nanomaterials is investigated under irradiation from a 60 – watt light bulb and photocatalytic activity is revealed to be due to the defects present in the crystal structure. The ZnS shows anti-bacterial action as both a bacteriostatic and bacteriocidal (88 % reduction in the amount of bacteria in 5 h) material and the methods of bacterial degradation on the ZnS is discussed. The anti-bacterial actions of these materials were also compared with commercial ZnS and Evonik-Degussa P-25. A detailed mechanism for the light absorption in the visible light region of the microwave prepared ZnS is proposed based on the luminescence spectroscopy.

**Key words**: Gram positive and gram negative bacteria, decontamination, visible light, XPS, solar, anti-microbial, ZnO, MRSA, green and energy efficient synthesis, anti-MRSA coatings, hospital acquired infections, photo-activity, emerging pollutants, $TiO_2$, Degussa P-25, Cell wall, disinfection, micro-biology, doping, band gap, *Escherichia coli*, *P. aerginosa,* methicillin-resistant *Staphylococcus Aureus* (MRSA).

## Introduction

Semiconductor photocatalysis has been shown to be an effective method of removing bacterial contamination, including *Escherichia coli* and *Staphylococcus aureus*, in hospitals and in the food industry.[1,2,3,4,5,6,7,8, 9] Zinc sulfide is an important II – VI semiconductor with a band gap of 3.8 eV.[10,11] Zinc sulfide has a great potential as an anti-bacterial semiconductor due to its relatively deeper conduction and valence bands compared to other semiconductors including $TiO_2$ and ZnO.[3,12] The comparatively negative positions of the conduction and valence bands of ZnS versus the NHE are located at -2.3 and +1.4 eV respectively. ZnS was one of the first semiconductors discovered and is of major interest in biomedical, electronic and photovoltaic devices due to its luminescent and catalytic properties.[13,14] It has been used as an important component in ultraviolet light emitting diodes,[15]

ACCEPTED MANUSCRIPT

injection lasers, flat panel displays,[16] solar cells,[17] and as a photocatalysts in many situations including a visible light activated water-splitting agent.[18]

ZnS has been synthesised in a number of ways including by precipitation method,[19] by spray pyrolysis,[20] by hydrothermal synthesis,[21] in solution,[22] using ligands,[23] with conjugated polymers,[24] using ultrasonic irradiation,[25] and employing microwave heating.[26] Microwave heating as a synthetic route is of great interest to materials scientists due to the capability of molecular level heating, which leads to homogenous and quick thermal reactions.[12] Microwave reactions have been used to prepare zinc sulfide but mostly these reactions require the use of pressurised containers made of quartz or Teflon.[26,27,28]

Semiconductor photocatalysts have been established as viable materials for the creation of self-sterilising environments and for the destruction of bacteria. This self-sterilisation is an important part of combating healthcare associated infections, which are a serious problem at every level of healthcare systems.[29] The incidence of healthcare associated infections, HAIs is growing worldwide, especially the number of cases of methicillin-resistant *Staphylococcus Aureus* (MRSA) in Europe,[30] North America,[31] and Australasia.[32,33] This increase is mainly due to the epidemics of highly transmissible clones.[34] Studies have shown that prevention techniques such as cleaning and the use of barriers, gloves and gowns, can be effective but that behavioural practices limit the effectiveness of these methods.[29] This creates the need for a material on location which will create a sterile environment, and semiconductor photocatalysts are an ideal solution.

When the semiconductor is irradiated by a photon of light, which matches or exceeds the bandgap energy, $E_g$, an electron is promoted from the valence band to the conduction band leaving a positive hole. The photo-generated hole in the valence band and the photo-generated electron in the conduction band can serve respectively as oxidation and reduction species.[35] Titanium dioxide, $TiO_2$, is the most widely studied semiconductor photocatalyst has been shown to kill bacteria on its surface,[36,37] has been incorporated into tiles to create self-cleaning surfaces[38] and into medical devices such as catheters.[39] The major advantage of using photocatalytic self-cleaning surfaces as a treatment for HAIs is that they operate in a passive manner, that is, without the need for chemicals or electric power, with only light, oxygen and water from the atmosphere required. The semiconductor surface is non-toxic and doesn't produce harmful by-products unlike some of the chemical reagents and cleaning products used as anti-bacterial agents.[40]

In the current investigation, a novel ambient pressure microwave assisted method of producing zinc sulfide nanocrystals, which is a quick and straightforward water based reaction, where the product is collected and used as is, is presented. The ZnS were characterized in detail and the anti-bacterial photocatalytic activities of these materials were tested against *E. coli*, and *S. aureus* and *P. aeruginosa.*

**Experimental**

**Preparation of Photocatalysts.**

In a typical synthesis, 200 mL of a 0.2 M zinc acetate dihydrate aqueous solution was added to 200 mL of a 0.2 M thiourea aqueous solution at room temperature. The beaker containing these solutions was placed in a MARS 5 microwave system and was irradiated at 600 W for 30 min under

ACCEPTED MANUSCRIPT

ambient pressure followed by a 5 min cool down period. After this time had elapsed the water had completely evaporated and a dry yellow powder was collected and characterized without further treatment. Samples prepared in this way are referred to by their irradiation power, *e.g.* ZnS450W – ZnS prepared at 450 W irradiation power.

**Characterisation**

The obtained ZnS was investigated using a combination of characteristic techniques including X-ray diffraction using a Siemens D 500 X-ray diffractometer with the diffraction angles scanning from 2θ = 20 – 80 °, using a Cu Kα radiation source. The diffuse absorbance spectra of the samples were measured by a UV-Vis-NIR Perkin Elmer Lambda 900 spectrometer between 800 and 200 nm with samples prepared in a KBr disc with a sample to KBr ratio of 1:10. Luminescence measurements were taken, with samples suspended in ethanol, by a Perkin Elmer LS 55 luminescence spectrometer. X-ray photoelectron spectrometer measurements were taken with a Thermo VG Scientific (East Grinstead, UK) Sigma Probe spectrometer using a monochromated A1 Kα X-ray source (hν = 1486.6 eV), which was used at 140 W and the area of analysis was approximately 500 µm in diameter. Infrared spectra were obtained using a Spectrum GX Infrared Spectrometer. SEM images were obtained using a Hitachi SU-70 FE-SEM. Surface area measurements were measured using a Quantachrome Nova 2200 with the samples degassed at 300 °C for 2 h.

**Anti-bacterial Testing – Agar Test Method**

Agar plates were prepared using Mueller Hinton agar plate mix. The loading of the powder samples in the agar was 1 % w/w. The agar was autoclaved at 120 °C prior to testing. After autoclaving the agar was heated to 100 °C and the photocatalyst was added to the molten agar with stirring and dispersed with sonication before being poured onto a plate. Each of the agar plates were inoculated with 50 µL *E. coli*, 50 µL of *S. aureus* and 50 µL of *P. aerginosa* suspension. Each of these suspensions was diluted to approximately 1000 colony forming units/mL (CFU/mL). Each plate was prepared in triplicate with one sample irradiated with a 60-watt light bulb (*i.e.* 3 plates irradiated) and one sample kept in the dark (*i.e.* three plates not irradiated). After irradiating the samples for 3 h with a 60-watt light bulb, both sets of plates were placed in an incubator at 37 °C for 24 h to allow any growth of the bacteria to take place.

**Anti-bacterial Testing – Suspension Test Method**

The suspension test measures the bactericidal effect of the photocatalyst and makes it possible to determine a time frame for the reduction in the bacterial colonies. The reference sample for these tests contained no catalyst or powder. The reference sample consisted of 4.5 mL multi-recovery diluents, MRD, and is inoculated with 500 µL of $10^6$ CFU/mL of bacteria. This gives a theoretical concentration of $10^5$ CFU/mL at time zero. The test sample contains 10 mg of photocatalyst suspended in 4.5 mL of MRD, which was then inoculated with 500 µL of $10^6$ CFU/mL of bacteria. The reference and sample were both exposed to light from a 60-watt light bulb for a total of 5 h. Samples from both were taken at 0 min, 30 min, 1 h, 3 h and 5 h.

**Results and Discussion**

The ZnS materials were prepared with microwave irradiation powers of 300, 450 and 600 watts. In each preparation, a solution of zinc precursor and a solution of sulfur precursor were mixed in an

ACCEPTED MANUSCRIPT

open beaker. The solutions were irradiated in the microwave and at the end of the irradiation period of dry yellow powder was collected and used for all tests without further washing or synthesis steps. The simplicity of this reaction allows for the production of nanomaterials in a cost effective and easily scalable reaction scheme.

The solutions undergo dielectric heating upon irradiation in the microwave and the following mechanism is believed to lead to the formation of the zinc sulfide.

$$H_2NCSNH_2 + 2H_2O \rightarrow CO_2 + H_2S + 2NH_3 \qquad (1)$$

$$Zn^{2+} + H_2S \rightarrow ZnS + H_2 \qquad (2)$$

In equation (1), the thiourea breaks down, at a temperature of 86 °C,[41] forming $H_2S$, which is the sulfur source. The $H_2S$ reacts with the liberated zinc ion from the zinc acetate solution, to form ZnS, equation (2). Zhao *et al.*[42] found that using zinc acetate as a precursor can lead to the ZnS agglomerating with the presence of the acetate. The acetate molecules interact with the surface of the ZnS crystals and cause the agglomeration. The agglomeration was investigated by SEM and infrared spectroscopy. The SEM images show agglomerations of ZnS particles, prepared at 450 W, of up to 100 nm in size. (**Figure 1).**

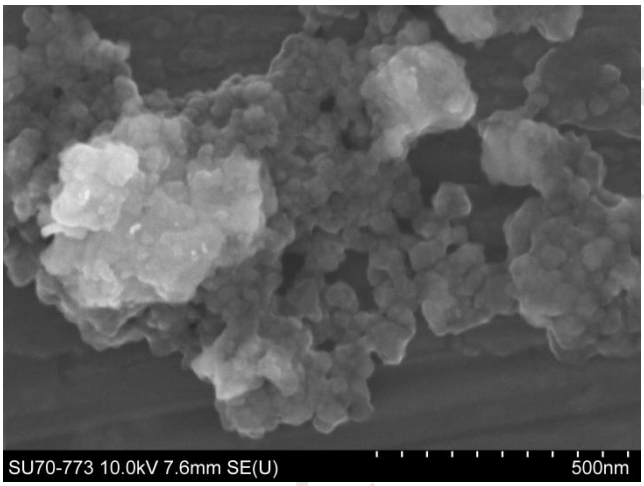



**Figure 1.** SEM image of ZnS prepared at 450 W.

X-ray diffraction was employed to determine the crystal structure of the as-prepared powders. The peak pattern show that all the samples prepared in the microwave are crystalline and sphalerite (cubic blende) phase zinc sulfide when compared to the International Centre for Diffraction Data (ICDD file #05-0566)[43,44] with the main peaks occurring at 28°, 47° and 56° corresponding to the (111), (220) and (311) planes respectively. (Figure 3) The broadening of these peaks, demonstrate that the materials are nanosized, and calculations based on the Scherrer equation show all particles to be 4 – 6 nm in size.

ACCEPTED MANUSCRIPT

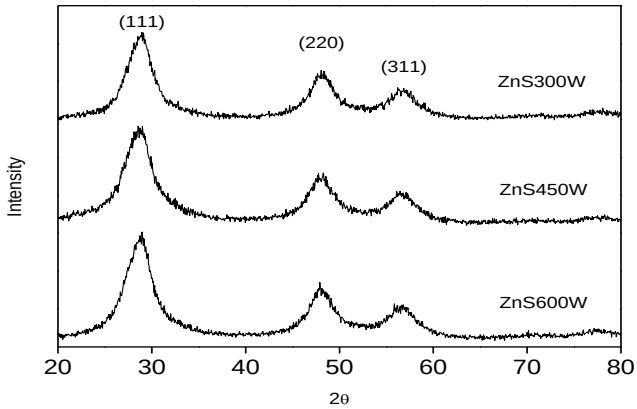
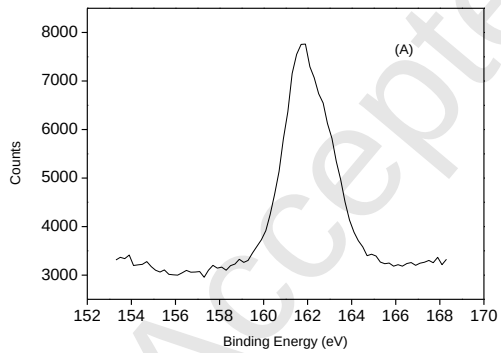
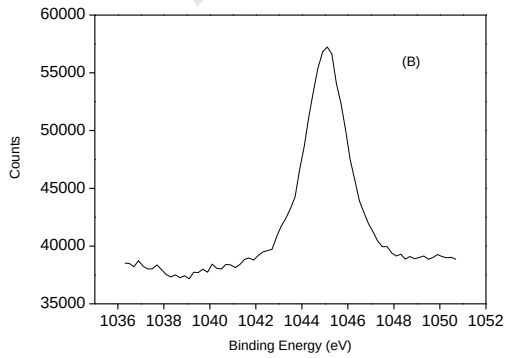



**Figure 2.** X-ray diffraction spectra of ZnS samples prepared via microwave irradiation.

X-ray photoelectron spectroscopy was used to further clarify the identification of the material. The XPS spectra of ZnS450W (prepared in the microwave at 450 W irradiation power) is typical of the XPS data for the materials prepared **(Figure 3)**. The XPS spectra for the S 2p, $Zn2p_{1/2}$ and $Zn2p_{3/2}$ locates the peaks at 161.9, 1045.1 and 1022.1 eV respectively, which confirm the material as zinc sulfide.[45,46] It should be noted that the S 2p spectra are composed of $2p_{3/2}$ and $2p_{1/2}$ peaks due to the spin-orbit splitting. Generally, the $S2p_{3/2}$ peak at lower binding energy is almost twice the intensity or area of the higher binding energy $S_2p_{1/2}$ peak. These peaks are only around1.1 eV apart and cannot be resolved from one another. The survey scan and more detailed scans indicate that there is no contaminants present on the surface of the material. The XPS spectra for the samples ZnS300W and ZnS600W are shown in supplementary information figures 1 and 2 respectively.




Accepted Manuscript

ACCEPTED MANUSCRIPT

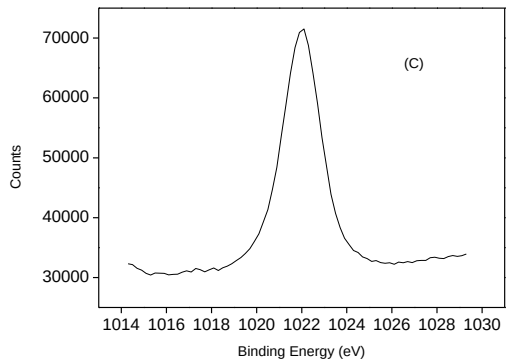
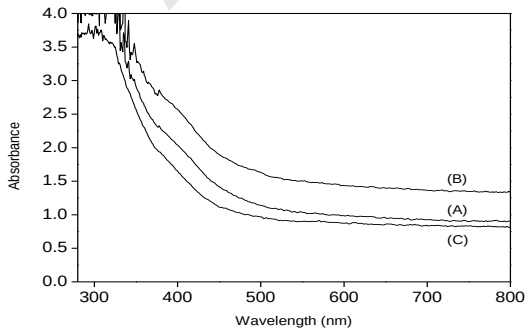



**Figure 3.** X-ray photoelectron spectra of the (a) S 2p (b) $Zn2p_{1/2}$ (c) $Zn2p_{3/2}$ peaks of ZnS450W.

The electronic properties of the as-prepared ZnS samples were investigated using a UV-Vis spectroscopy (**Figure 4)**. The spectra for ZnS300W, ZnS450W and ZnS600W are shown, and it can be seen that the absorption edges from 320 to 340 nm which equate to band gaps of 3.8 to 4.0 eV. The band gap is greater than that of the bulk ZnS (3.7 eV) due to the quantum confinement effect associated with nanosized materials.[47] For these samples there is an absorption shoulder which extends to 420 nm and along with the yellow colour of the powder shows visible light absorption (λ ≥ 390 nm). Luminescence spectroscopy offers an insight into the mechanism of the visible light absorption. The luminescence spectrum of ZnS450W was obtained upon excitation at 275 nm (**Figure 5)**. The peaks at 385, 427 and 489 nm correspond to three point defects within the ZnS crystal with the peak at 385 nm related to the interstitial Zn, the peak at 427 nm being from S vacancies and the peak at 489 nm relating to Zn vacancies.[19,46] These peaks and their related defects explain the visible light absorption as the levels of the defects act as donor levels within the band gap. Peng *et al.*[19] and others[48,49] showed that these defect sites can act as inter-band donor levels and reduced the energy required to promote an electron to the conduction band and leave a hole in the valence band, thus inducing photocatalysis. The luminescence spectra for the ZnS300W and ZnS600W are shown in supplementary information figures 3 and 4 respectively. Based on these luminescence and absorption spectroscopy results, a schematic is proposed for the mechanism of visible light absorption. **(Figure 6)** The donor levels, provided by the Zn and S vacancies, act as stepping stones for electrons travelling from the valence band to the conduction band. The donor levels energy gaps between the valance band and conduction bands are 2.5 and 2.9 eV respectively, which fall inside the region for visible light absorption. The main peak in the luminescence spectra at 385 nm which dominates the radiative decay lies inside the UV region of the spectrum.



ACCEPTED MANUSCRIPT

**Figure 4.** Diffuse Absorbance spectra of (A) ZnS300W (B) ZnS450W and (C) ZnS600W.

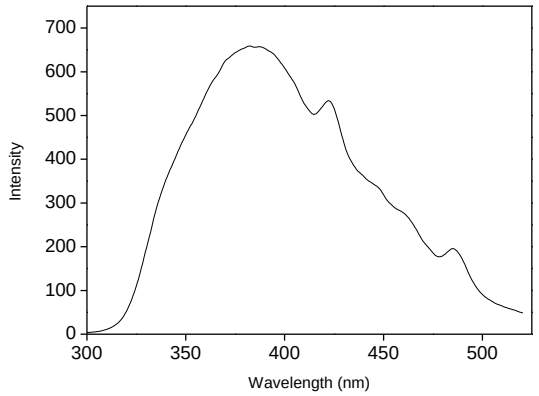



**Figure 5.** Luminescence spectra of ZnS450W excited at 275 nm.

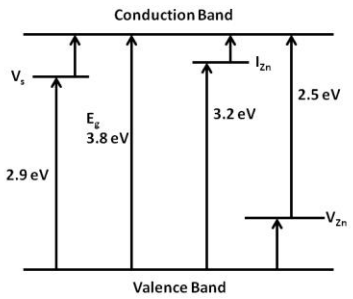



**Figure 6.** Schematic energy level diagram of ZnS

The microwave prepared ZnS was tested for its anti-bacterial activity under irradiation from a visible light source. The ZnS (**Table 2)** sample tested was ZnS300W and three references were used, a blank, Degussa P25 $TiO_2$, a commercial photocatalyst standard and Sigma-Aldrich ZnS. Initial testing involved using 1% w/w of photocatalyst in agar and it was found that the ZnS produced in the microwave had anti-bacterial properties in both light and dark conditions. The reference samples didn't show any anti-bacterial properties as can be seen in Table 2. The agar plates were inoculated with both *E. coli* and *S. aureus* and it was found that the ZnS produced in the microwave prevented the growth of colonies of both of these bacteria. The reason for this is believed to be due to the small particle size of the ZnS and the high concentration of the photocatalyst in the agar, affecting the growth of the colonies. **Figure 7(a)** shows agar plates with no photocatalyst, exposed to light and **(b)** the blank agar plate kept in the dark. Both plates show growth of *E. coli* and *S. aureus* bacterial colonies. **Figure 7(c)** shows the plate loaded with 1% w/w of the ZnS300W photocatalyst to light and **(d)** the agar loaded with the photocatalyst kept in the dark. Both plates, Figure 7 (c) and (d), show that no growth of the bacterial colonies of *E. coli* or *S. aureus* occurred. To determine the minimum inhibition concentration of ZnS required, to effectively prevent the growth of the bacterial colonies, plates containing 0, 0.25, 0.75 % w/w ZnS300W were inoculated with *E. coli*, *S. aureus* and *P.*

ACCEPTED MANUSCRIPT

*aerginosa*. Each plate was exposed to light from a 60 watt bulb for 3 h with a control plate for each concentration kept in the dark for the same time. After the irradiation time, the plates were kept overnight in an incubator at 37 °C. The plate containing no catalyst showed growth of all three bacteria in both light and dark conditions. At 0.25 % w/w loading, in the dark conditions both the *S. aureus* and *E. coli* grew colonies, however on the plates exposed to light the growth of *S. aureus* was partially inhibited while the growth of *E. coli* colonies was fully inhibited. This is an indication of photocatalytic action against the bacteria by the microwave synthesized ZnS. At 0.75 % w/w loading inhibition of *E. coli* and *S. aureus* bacteria was observed in both light and dark conditions, however the growth of *P. aerginosa*, a species resistant to many anti-bacterial agents, was not affected.

These tests showed that the ZnS acts as a bacteriostatic material preventing the growth of the test microbes in both light and dark conditions at concentrations above 0.75 % w/w and that photocatalytic action takes place when the samples are exposed to light at concentration at low as 0.25 % w/w.

To observe the bactericidal effects of the material a suspension test was designed. The reference samples contained no catalyst and comprised of 4.5 mL of multi-recovery dilutent (MRD) and were inoculated with 500 µL of $10^6$ CFU/mL of bacteria. The test samples contained 10 mg of ZnS300W suspended in 4.5 mL of MRD, which was inoculated with $10^6$ CFU/mL of bacteria. This gave a concentration of $10^5$ CFU/mL of bacteria in each sample. The reference and test samples were exposed to light from the 60 watt bulb for 5 h with samples taken at 0 min, 30 min, 1 h, 3 h and 5 h and the concentration of bacteria was measured at each time point. **Table 3** shows the concentration of bacteria at each time point for both the reference and test sample and **figure 8** shows a graph of the relative concentrations of bacteria. The test showed that there was an 88 % reduction in the amount of bacteria in 5 h and showed the material to be bactericidal.

The study in this paper where the concentration of ZnS is varied and previous studies have shown that the anti-bacterial effect of the semiconductor is concentration dependant.[1,50] The increase in ZnS concentration leads to a decrease in the bacterial concentration, and in this study when the concentration was above 0.75 % w/w of the agar, no bacterial colonies of *S. aureus* or *E. coli* were observed. It has been found that below the concentration threshold for which the ZnS is harmful to the bacteria that the number of bacterial colonies actually increases due to the fact that bacteria can metabolise $Zn^{2+}$ at low concentrations.[1] A number of factors may play a role in the antibacterial effect of ZnS. *S. aureus* has an efflux mechanism for $Zn^{2+}$ which suggests that free cations on the surface of the material generated by the photo dissociation of ZnS is toxic to the bacteria.[51] Oxidative species such as superoxide anions and singlet oxygen can be generated on the surface of the photocatalyst after irradiation with light[52] will also assist in the degradation of the cell membrane of the bacteria which leads to loss of cellular components and induces cell death.[4,53] The small size of the nanoparticle as shown by the broadening of the XRD peaks a larger surface area for contact with the bacteria also increases the antibacterial activity of the ZnS.

ACCEPTED MANUSCRIPT

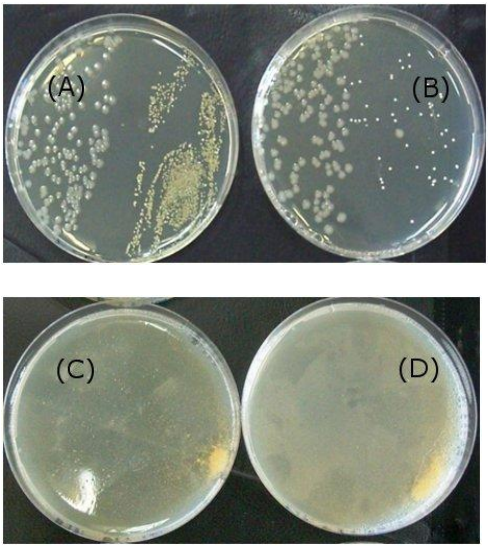



**Figure 7.** (a) Blank agar exposed to light (L+), (b) blank agar kept in the dark (L-), (c) agar plate loaded with 1% w/w ZnS exposed to light (L+) and (d) the agar plate loaded with 1% w/w kept in the dark (L+).

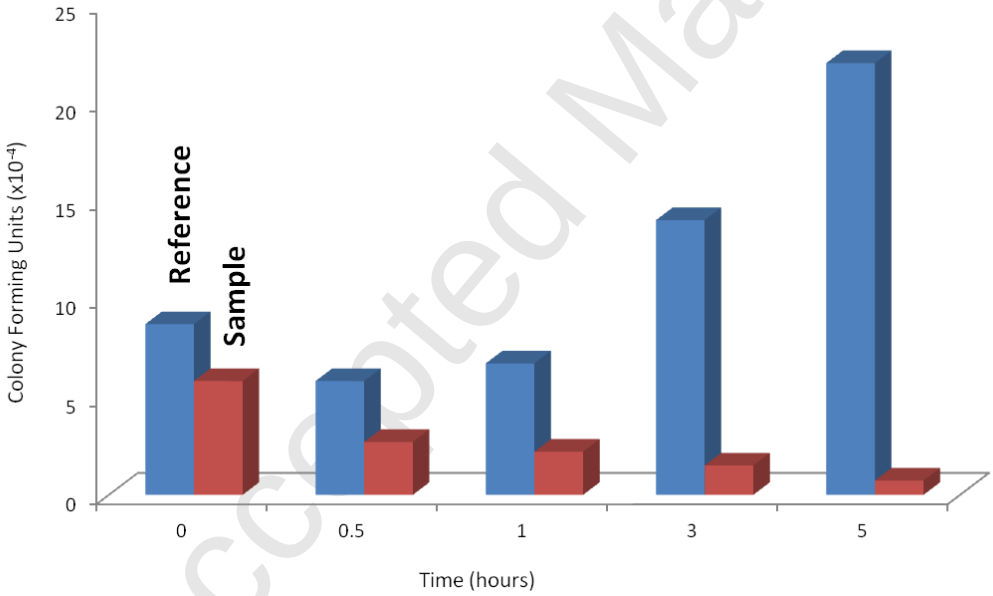



**Figure 8.** Bacterial concentration of reference and ZnS300W sample to test the degradation rate of the photocatalyst.

**Table 2.** Colony growth of *S. aureus* and *E. coli* on agar plates loaded with 1 % w/w photocatalyst.

| **Photocatalyst** | **Light exposure** | | **No Light Exposure** | |
|---|---|---|---|---|
| | ***S. aureus*** | ***E. coli*** | ***S. aureus*** | ***E. coli*** |
| **Blank** | Growth | Growth | Growth | Growth |

Accepted Manuscript

ACCEPTED MANUSCRIPT

| | | | | |
|---|---|---|---|---|
| **Degussa-Evonik P25 $TiO_2$** | Growth | Growth | Growth | Growth |
| **Commercial ZnS (Sigma-Aldrich)** | Growth | Growth | Growth | Growth |
| **ZnS300W** | No Growth | No Growth | No Growth | No Growth |

**Table 3**. Concentrations of the bacterial suspension test for the reference and test samples.

| **Time** | **0 h** | **0.5 h** | **1 h** | **3 h** | **5 h** |
|---|---|---|---|---|---|
| **Reference CFU count** | $8.7 \times 10^5$ | $5.8 \times 10^5$ | $6.7 \times 10^5$ | $1.4 \times 10^6$ | $2.2 \times 10^6$ |
| **Sample CFU count** | $5.8 \times 10^5$ | $2.7 \times 10^5$ | $2.2 \times 10^5$ | $1.5 \times 10^5$ | $7.3 \times 10^4$ |

**Conclusions**

In summary, a method of synthesizing ZnS nanocrystals which is straight-forward, cost-effective and easily scalable has been created. The material produced is nanosized, crystalline and pure zinc sulfide with particles that are 4 nm in size. The material shows excellent antibacterial activity under irradiation from indoor light, conditions under which commercially available photocatalysts are inactive. The material is shown to be both photocatalytically bacteriostatic and bacteriocidal, preventing bacteria from growing on the surface and killing bacteria that are present. 88 % reduction in the concentration of bacteria was observed within 5 h of visible light irradiation. It was also observed that the increase in ZnS nanomaterial concentration leads to a decrease in the bacterial colonies and when the concentration was above 0.75 % w/w of the agar, no bacterial colonies of *S. aureus* or *E. coli* were detected. It is envisaged that the small crystalline size and the high surface area of the ZnS nanoparticles are the major reasons for the significantly higher anti-bacterial activity compared to the commercial control samples. Luminescence spectroscopy reveals the possible mechanism of the indoor light activity with Zn and S vacancy defects acting as donor levels (2.5 and 2.9 eV) in the band gap allowing light in the visible region, up to 500 nm, to cause electron transfer from the valance band to the conduction band.

Accepted Manuscript

ACCEPTED MANUSCRIPT

**Acknowledgements**

The authors wish to thank Science Foundation Ireland (SFI grant number 10/US/I1822) for supporting this investigation under the US-Ireland R&D partnership programme, CREST and Focas for lab and office facilities and equipment. The authors would also like to thank Dr. John Colreavy for reviewing the paper and providing valuable comments and Aisling Kirwan for providing the scanning electron microscopy images. The authors would like to acknowledge Enterprise Ireland Nanovigil (CFTD/06/IT/326 and ARE/2008/0005) for financial support.

[1] M. Pelaez, N. T. Nolan, S. C. Pillai, M. K. Seery, P. Falaras, A. G. Kontos, P. S.M. Dunlop, J. W.J. Hamilton, J. Byrne, K. O'Shea, M. H. Entezari, D. D. Dionysiou, Appl. Catal. B, 125 (2012) 331.

[2] A. Kubacka, M. Ferrer, A. Martinez – Anas, M. Fernández – Garcia, Appl. Catal. B 84 (2008) 87
[3] X. Hu, C. Hu, J. Qu Appl. Catal. B 69 (2006) 17
[4] A. G. Rincón, C. Pulgarin Appl. Catal. B 49 (2004) 99
[5] A. K. Benabbou, Z. Derriche, C. Felix, P. Lejeune, C. Guillard Appl. Catal. B 76 (2007) 257
[6] P. S. M. Dunlop, M. Ciavola, L. Rizzo, J. A. Byrne, Chemosphere 85 (2011) 1160
[7] P. K. J. Robertson, J. M. C. Robertson, D. W. Bahnemann, J. Hazard. Mater. 211 (2012) 161
[8] J. A. Byrne, P. A. Fernandez-Ibanez, P. S. M. Dunlop, D. M. A. Alrousan, J. W. J. Hamilton, Int. J. Photoenergy, doi: 10.1155/2011/798051
[9] O. M. Alfano, D. W. Bahnemann, A. E. Cassano, R. Dillert, R. Goslich, Catal. Today, 58 (2000) 199
[10] X. Fang, Y. Bando, M. Liao, U. K. Gautam, C. Zhi, B. Dierre, B. Liu, T. Zhai, T. Sekiguchi, Y. Koide, D. Golberg Adv. Mater. 21 (2009) 2034
[11] S. H. Yu, M. Yoshimura Adv. Mater. 14 (2002) 296
[12] Y. Wada, H. Yin, S. Yanagida Catal. Surv. Jpn. 5 (2002) 127
[13] X. Fang, Y. Bando, G. Shen, C. Ye, U. K. Gautam, M. Pedro, F. J. Costa, C. Zhi, C. Tang, D. Golberg Adv. Mater. 19 (2007) 2593
[14] C. Feigl, S. P. Russo, A. S. Barbard, J. Mater. Chem. 20 (2010) 4971
[15] D. D. Fanfair, B. A. Korgel Cryst Growth Des. 8 (2008) 3246
[16] L. Zepang, L. Bingbing, L. Xianglin, Y. Shidan, L. Wang, H. Yuanyuan, Z. Yonggang, M. Mingguang, L. Quanjun, Z. Bo, C. Tian, Z. Guantian, W. Guorui, L. Yichun Nanotechnology 18 (2008) 255602
[17] H. Zhang, J. F. Banfield J. Phys. Chem C. 113 (2009) 9681
[18] J. H. Bang, R. J. Helmich, K. S. Suslich Adv. Mater. 20 (2008) 2599
[19] W. Q. Peng, G. W. Cong, S. C. Qu, Z. G. Wang Opt. Mater. 29 (2006) 313
[20] E. Bacaksiz, O. Görür, M. Tomakin, E. Yanmaz, M. Altunbas Mater. Lett. 61 (2007) 5239
[21] C. Jiang, W. Zhang, G. Zou, W. Yu, Y. Qian Mater. Chem. Phys. 103 (2007) 103
[22] G. Feng, C. Z. Li, S. F. Wang, M. K. Lü Langmuir 22 (2006) 1329
[23] Y. Li, X. Li, C. Yang, Y. Li J. Phys. Chem. B. 108 (2004) 16002
[24] Y. Gou, Z. Su, Z. Xue, Mater. Res. Bul. 39 (2004) 2203
[25] G. Z. Wang, B. Y. Geng, X. M. Huang, Y. W. Wang, G. H. Li, L. D. Zhang Appl. Phys. A. 77 (2003) 933
[26] T. Rath Inorg. Chem 47 (2008) 3014
[27] W. Jian, J. Zhuang, W. Yang, Y. Bai J. Luminescence 126 (2007) 735
[28] H. Yang, C. Huang, Y. Su, A. Tang J. Alloy and Comp. 402 (2005) 274
[29] S. W. Aboelela, P. W. Stone, E. L. Larson J. Hosp. Infect. 66 (2007) 101
[30] European Antimicrobial Resistance Surveillance System Annual Report 2003
[31] D. J. Diekama, B. J. Bootsmillar, T. E. Vaughn Clin. Infect. Dis. 38 (2004) 521

Accepted Manuscript

ACCEPTED MANUSCRIPT

[32] B. P. Howden, P. B. Ward, P. G. P. Charles Clin. Infect. Dis. 38 (2004) 52
[33] S. Kim, W. Park, K. Lee Clin. Infect. Dis. 37 (2003) 794
[34] I. M. Gould J. Hosp. Infect. 61 (2005) 277
[35] X. Li, S. Ouvang, S. Kikugawa, J. Ye J. Appl. Catal. A 334 (2008) 51
[36] A. Fujishima, K. Hashimoto, T. Wantanbe $TiO_2$ Photocatalysis: Fundamentals and Applications, 1999, BKC Inc. Toyko
[37] A. Fujishima, T. N. Rao, D. A. Tryk J. Photochem. Photobiol. C. 1 (2001) 1
[38] T. N. Rao, A. Fujishima, D. A. Tryk, A. J. Bard, M. Stratman, S. Licht (eds) Encyclopedia of Electrochemistry in Semiconductor Electrodes and Photoelectrochemistry, vol 6 (2002) Weinheim
[39] K. Ohko, Y. Utsumi, C. Miwa, T. Tatsuma, K. Kobayakawa, Y. Saton, Y. Kubota, A. Fukishima J. Biomed. Mater. Res. 58 (2001) 97
[40] A. Fujishima, X. Zhang, D. A. Tryk Surf. Sci. Rep. 63 (2008) 515
[41] H. Shao, X. Oian, Z. Zhu J. Solid State Chem. 178 (2005) 3522
[42] Y. Zhao, J. M. Hong, J. J. Zhu J. Crys. Grow. 270 (2004) 438
[43] K. P. Subhendu, A. Dalta, S. Chaudhari, Chem. Phys. Lett. 440 (2007) 235
[44] H. Zhang, Q. Li, Nanotechnology 17 (2006) 3984
[45] C. M. Huang, K. W. Cheng, Y. R. Jhan, T. W. Chung, Thin Solid Films 515 (2007) 7935
[46] M. V. Limaye, S. Gokhale, S. A. Acharya, S. K. Kulkarni, Nanotechnology 19 (2008) 415602
[47] Y. Zhao, J. M. Hong, J. J. Zhu, J. Crystal Growth 270 (2004) 438
[48] M. Scocioreanu, M. Baibarac, I Baltog, I Pasuk, T. Velula J. Solid State Chem. 186 (2012) 217
[49] S. K. Mehta, S. Kumar J. Lumin. 120 (2010) 2377
[50] R. Brayer, R. Ferrari – Iliou, N. Brivois, S. Djediat, M. F. Benedetti, F. Fievet Nano Lett. 6 (2006) 866
[51] K. M. Reddy, K. Feris, J. Bell, D. G. Wingett, C. Hanley, A. Punnoose Appl. Phys. Lett. 90 (2007) 213902
[52] M. Pelaez, P. Falcras, V. Likodimos, A. G. Kontos, A. A. De la Cruz, K. O' Shea, D. D. Dionysiou Appl. Catal. B 99 (2010) 378
[53] J. Kiwi, V. Nadtochenko, *Langmuir* 21 (2005) 4631.

SYNOPSIS TOC (Word Style "SN_Synopsis_TOC"). If you are submitting your paper to a journal that requires a synopsis graphic and/or synopsis paragraph, see the Guide, Notes, Notice, or Instructions for Authors that appear in each publication's first issue of the year and the journal's homepage for a description of what needs to be provided and for the size requirements of the artwork.

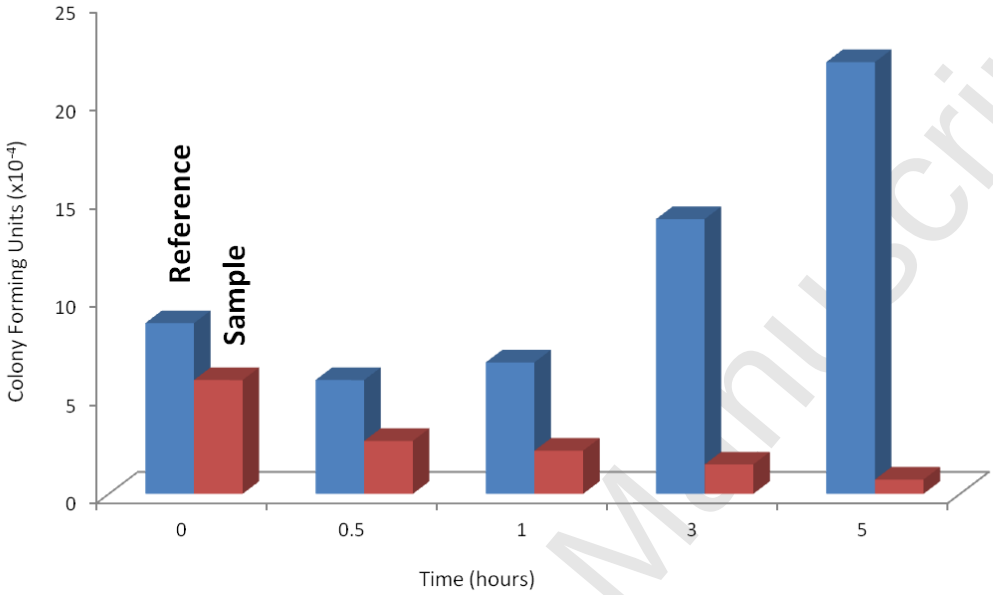
Accepted Manuscript

 ACCEPTED MANUSCRIPT

Graphical Abstract

Anti-Bacterial Activity of Indoor Light Activated Photocatalysts

Damian W. Synnotta, Michael K. Seery, Steven J. Hinder, Georg Michlits, and Suresh C. Pillai




Degradation rate of the bacteria in the presence of the photocatalyst.

Accepted Manuscript

ACCEPTED MANUSCRIPT

Research Highlights

- Nanocrystalline photocatalysts, prepared under ambient conditions using a microwave assisted synthesis
- These nanomaterials show indoor light photocatalytic activity for the degradation of *S. aureus* and *E. coli*.
- A novel energy efficient method (green synthesis) to produce anti-bacterial zinc sulphide.
- The mechanism for the indoor light antibacterial photocatalysis is proposed.

Accepted Manuscript